

## Solutions

### Independent Practice

4. What is the molarity of a solution composed of 33.2g of potassium iodide, KI, dissolved in enough water to make 0.125L of solution?

1.60 M KI



5. How many moles of  $\text{H}_2\text{SO}_4$  are present in 0.500L of a 0.150M  $\text{H}_2\text{SO}_4$  solution?

0.075 moles  $\text{H}_2\text{SO}_4$

6. What volume of 4.00M NaCl is needed for a reaction that requires 146g of NaCl?

0.625 L

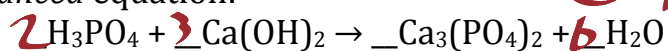
7. What is the molarity of a 2.0L solution that is made from 14.6g of NaCl?

0.125 M NaCl

8. What volume of a 0.500M solution of HBr is needed for a reaction that requires 20.2g of HBr?

0.500 L

9-10. Consider the following *unbalanced* equation:



~~Challenge~~ Challenge

9. What mass of calcium phosphate results if 750mL of 6.00M  $\text{H}_3\text{PO}_4$  reacts completely according to the equation?

10. What mass of water results?

9) 697.

10) 243